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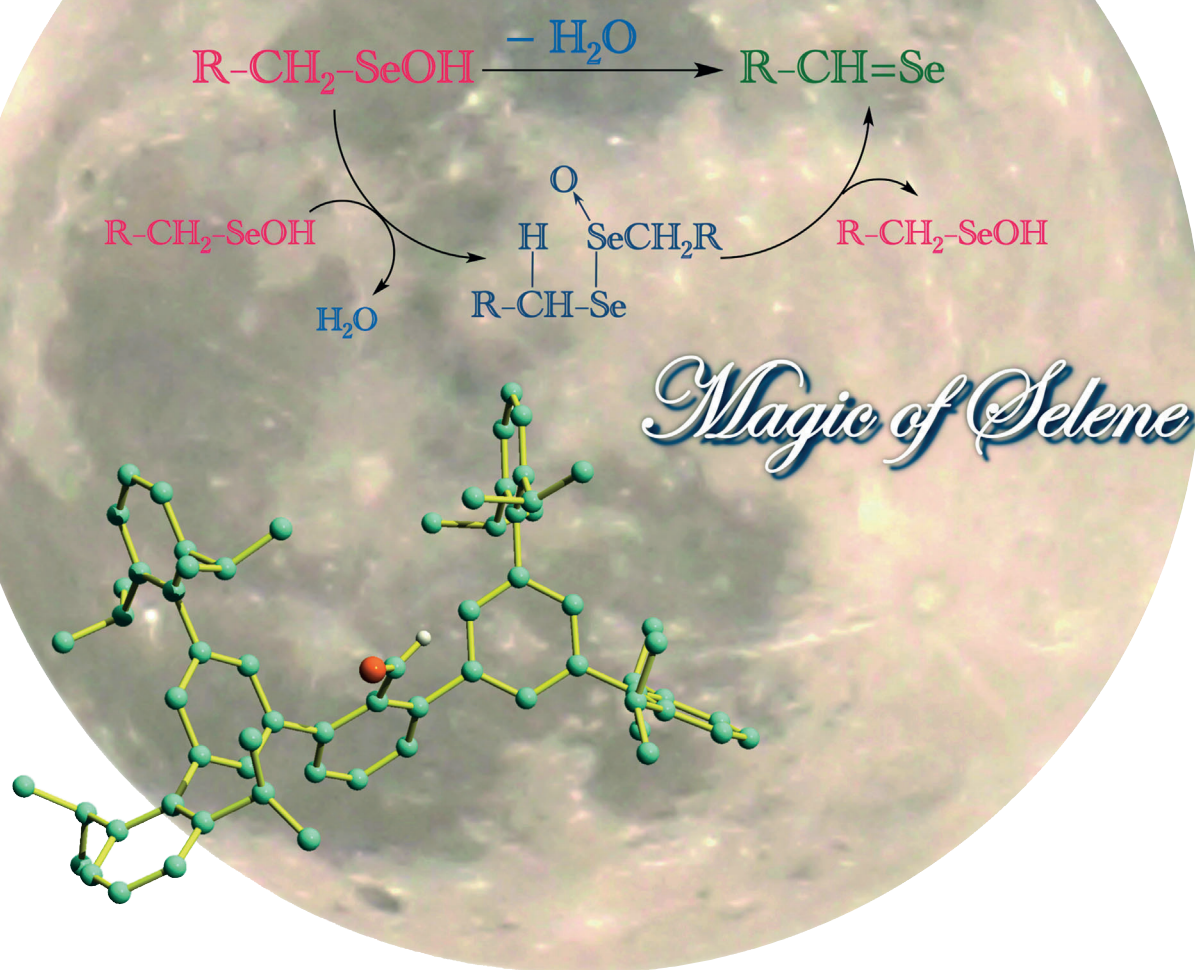
# Angewandte Chemie

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## Depriving a selenenic acid ( $\text{RCH}_2\text{-SeOH}$ ) ...

... of water results in its transformation into a selenoaldehyde ( $\text{RCH=Se}$ ). In their Communication on page 901 ff., K. Goto and co-workers investigate the  $\beta$ -dehydration reaction to form a C=Se bond. The selenenic acid self-catalyzes this transformation through a dehydrative condensation (to form a selenoseleninate intermediate) and a subsequent  $\beta$ -elimination. A dendrimer-type cradle protects both the selenenic acid and the selenoaldehyde, which are otherwise evanescent.

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